Module 15 Exploring Creation With Chemistry by Dr. Jay L. Wile
Chemical Equilibrium- The point at which both the forward and reverse reactions in a chemical equation have equal reaction rates. When this occurs, the amounts of each substance in the chemical reaction will not change, despite the fact that both reactions still proceed.(524)
Le Chatelier's Principle- When a stress (such as a change in concentration, pressure, or temperature) is applied to an equilibrium, the reaction will shift in a way that relieves the stress and restores equilibrium.(537)
Acid Ionization Reaction- The reaction in which an H+ separates from the acid molecule so that it can be donated in another reaction.(544)
Acid Ionization Constant- The equilibrium constant for an acid's ionization reaction.(545)

Topics:
- ☐ Introduction, The Definition of Chemical Equilibrium
- ☐ The Equilibrium Constant
- ☐ A Few more Details Concerning the Equilibrium Constant
- ☐ Using the Equilibrium Constant to Predict the Progress of a Reaction
- ☐ Le Chatelier's Principle
- ☐ Acid / Base Equilibria
- ☐ The pH Scale
- ☐ Acid Rain

Notes:

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Oxidation Number- The charge that an atom in a molecule would develop if the most electronegative atoms in the molecule took the shared electrons from the less electronegative atoms.(557)
Oxidation- The process by which an atom loses an electron.(563)
Reduction- The process by which an atom gains an electron.(563)

Topics:
- ☐ Introduction, Oxidation Numbers
- ☐ Oxidation and Reduction
- ☐ Recognizing Reduction / Oxidation Reactions, Experiment 16.1
- ☐ How Batteries Work
- ☐ Real Batteries
- ☐ Corrosion, A Few final Words

Notes: